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ANSWER Answer

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1. _____ The
molar mass of a
gas at STP is
the density of
that gas (a)
multiplied by
the mass of 1
mol. (c)
multiplied by
22.4 L. (b)
divided by the
mass of 1 mol.

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(d) divided by
22.4 L. 2.
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primary
components of
air are carbon

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dioxide,
nitrogen, and
oxygen. In a
sample

containing a
mixture of only
these gases at
exactly 1 atm,
the partial
pressures of
carbon dioxide
and nitrogen are
given as P_{CO_2}
 $= 0.285$ torr and

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$P_{N_2} = 593.525$

torr. What is
the partial
pressure of
oxygen?

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continued 5.

Convert a

pressure of

0.200 atm to the

following units:

152 a. mm Hg

20.3 b. kPa 6.

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When an explosive like TNT is detonated, a mixture of gases at high temperature is created. Suppose that gas X has a pressure of 50 atm, gas Y has a pressure of 20 atm, and gas Z has a pressure

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of 10 atm. 80
atm a. What is
the total ...
Answer Key

*mc06se cFMsr i-
vi*

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ANSWER Answer
the following
questions in the
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1. List the following gases in order of increasing rate of effusion.

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ANSWER Answer

the following
questions in the
space provided.

1. _____ The
molar mass of a
gas at STP is
the density of
that gas (a)

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Chemistry by
the mass of 1
mol. (c)
multiplied by
22.4 L. (b)
divided by the
mass of 1 mol.
(d) divided ...

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Volumes and the
Ideal Gas Law In
the space

provided, write
the letter of
the term or
phrase that best
completes each

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Chemistry or best
answers each
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question. _____

Answer Key
1. At the same
temperature and
pressure,
balloons of
equal volume
always contain

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